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Solution: $p_{abs} = p_{atm} + p_{gage} = 101.3kPa + 20 \text{ in. H}_2\text{O} \cdot 1.248 \cdot 8Pa = 110.26kPa$ Ideal gas law: $\rho = p / RT = 110.26kPa / (1 \text{ kgK} \cdot 2077J / 1.293 \cdot 2K \cdot 1000Pa \cdot 1kPa / J \cdot Nm \cdot N \cdot Pam^2) = 0.181 \text{ kg/m}^3$ b.) Situation: A sphere of 93 mm diameter contains an ideal gas. $T = 20 \text{ C} = 293.2K$ Find: Calculate the density of argon at a vacuum pressure of 8.8 psi. Properties:

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