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The number  $6.022 \times 10^{23}$ , called Avogadro's number after the 19th-century chemist Amedeo Avogadro, is the number we use in chemistry to represent macroscopic amounts of atoms and molecules. Thus, if we have  $6.022 \times 10^{23}$  Oxygen atoms, we say we have 1 mol of Oxygen atoms.

### CH105: Chapter 1 - Measurements in Chemistry - Chemistry

Modern Chemistry 6 The Periodic Law CHAPTER 5 REVIEW The Periodic Law SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. \_\_\_\_ When an electron is added to a neutral atom, energy is (a) always absorbed.

### Modern Chemistry Chapter 5 Review Answers Section 1

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