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Chapter 9 - Stoichiometry 9-1  
Introduction to Stoichiometry

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2  
Composition Stoichiometry - deals with mass relationships of elements in compounds  
Reaction Stoichiometry - Involves mass relationships between reactants and products in a chemical reaction  
I. Reaction Stoichiometry Problems  
A. Four problem Types, One Common Solution

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## 2 Chapter 9 - Stoichiometry

CHAPTER 9 REVIEW Stoichiometry

SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided.

1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield.

2. 6.0 mol of  $N_2$  are mixed with 12.0 mol of H



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2 comes from the Greek ... SECTION 9-1 OBJECTIVES Define stoichiometry. Describe the importance of the mole ratio in stoichiometric calculations. Write a mole ratio relating ... uses in a calculation. The number of significant figures in the answer is

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Example 2: Solution

Stoichiometry-Volume to Volume

Conversion. A student takes a precisely measured sample, called an aliquot, of 10.00 mL of a solution of  $\text{FeCl}_3$ . The student carefully adds 0.1074 M  $\text{Na}_2\text{C}_2\text{O}_4$  until all the  $\text{Fe}^{3+}(\text{aq})$  has precipitated as  $\text{Fe}_2(\text{C}_2\text{O}_4)_3(\text{s})$ .



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2 Using a precisely measured tube called a burette, the student finds that 9.04 mL of the  $\text{Na}_2\text{C}_2\text{O}_4$  solution was ...

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22 g and the theoretical yield is 25 ...

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Composition stoichiometry deals with the mass relationships of elements in compounds. Reaction stoichiometry involves the mass relationships between

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2 reactants and products in a chemical reaction. Reaction stoichiometry, the subject of this chapter, is based on chemical equations and the law of conservation of mass. All reaction stoichiometry

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2 Chapter 9 Stoichiometry Definition •  
Stoichiometry -Relationship between  
quantities • Composition stoichiometry  
-The mass relationships of elements in  
compounds (Ch 7.3) • Reaction  
stoichiometry -The mass relationships  
between reactants and products in a  
chemical reaction Section 1 Introduction  
to

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## 2 Chapter 9 Review Stoichiometry Answers Section 1

366. Stoichiometry. BIG Idea Mass relationships in chemical reactions confirm the law of conservation of mass.

11.1 Defining Stoichiometry MAIN Idea

The amount of each reactant present at the start of a chemical reaction determines how much product can form.

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